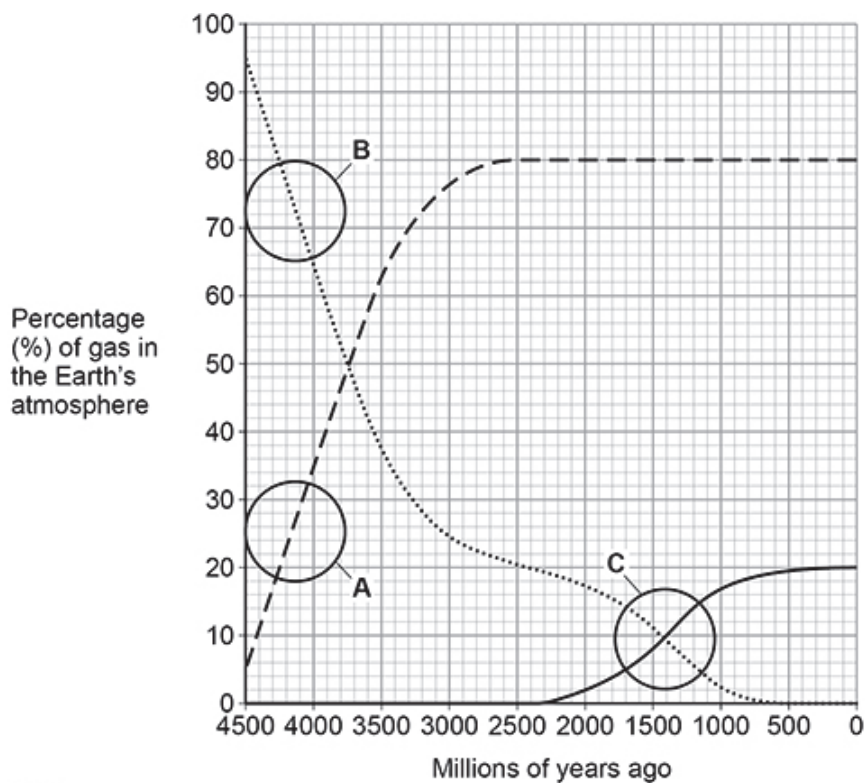


All questions are for both separate science and combined science students

Q1.

This question is about the chemistry of the Earth's atmosphere.

The figure below shows how the percentages of gases in the Earth's atmosphere may have changed since the atmosphere was formed.

**Key**

..... Carbon dioxide

- - - Nitrogen

— Oxygen

- (a) Explain the change in the percentage of gas in the region labelled **A** on the figure above.

(2)

- (b) Explain the change in the percentage of gas in the region labelled **B** on the figure above.

(2)

- (c) Compare the changes in the percentages of gases in the region labelled **C** on the figure above.

(2)

- (d) What process caused the changes in the percentages of gases in the region labelled **C** on the figure above?

(1)

- (e) Natural gas is a fossil fuel.

Describe how deposits of natural gas were formed.

(3)

(Total 10 marks)

Questions are for both separate science and combined science students unless indicated in the question

Q1.

This question is about burning fuels in central heating boilers.

In the future, gas central heating boilers may burn hydrogen rather than natural gas.

The table below shows information about these fuels when 1 dm³ of the fuel is burned in a central heating boiler.

	Fuel	
	Hydrogen	Natural gas
Energy released in kJ	11.9	37.1
Mass of carbon dioxide produced in grams	0.00	1.83
Mass of water vapour produced in grams	0.75	1.50
Mass of oxides of nitrogen produced in grams	6.6×10^{-4}	4.9×10^{-4}

- (a) Explain how oxides of nitrogen are produced when burning fuels.

(2)

- (b) Explain **one** positive impact on the environment of burning hydrogen rather than natural gas as a fuel.

Use the table above.

(2)

- (c) Explain **one** negative impact on the environment of burning hydrogen rather than natural gas as a fuel.

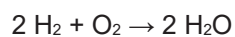
Use the table above.

(2)

- (d) Air is 20% oxygen.

Calculate the volume of air needed to provide enough oxygen to react with 3.50 dm³ of hydrogen gas. **(chemistry only) (HT only)**

The equation for the reaction is



Volume of air = _____ dm³

(3)

- (e) Central heating boilers can also burn kerosene.

Kerosene is produced from crude oil in a fractionating column using fractional distillation.

In the first step, crude oil is heated and hydrocarbon vapours are formed.

Explain how kerosene is produced from these hydrocarbon vapours.

(3)

(Total 12 marks)

Q2.

This question is about greenhouse gases and climate change.

Carbon dioxide and methane are greenhouse gases.

(a) Which of the following is also a greenhouse gas?

Tick (✓) **one** box.

Chlorine

Nitrogen

Oxygen

Water vapour

(1)

In the past 50 years, there has been an increase in:

- the world population
- the concentration of carbon dioxide in the atmosphere
- the concentration of methane in the atmosphere
- the mean temperature of the atmosphere at the Earth's surface.

Most scientists think this information can be used to explain climate change.

(b) Explain why the increase in world population may have caused the increase in the concentration of carbon dioxide in the atmosphere.

(2)

(c) Explain why the increase in world population may have caused the increase in the concentration of methane in the atmosphere.

(2)

- (d) Describe **two** potential effects of the increase in the mean temperature of the atmosphere at the Earth's surface.

1 _____

2 _____

(2)

- (e) The mean temperature of the atmosphere at the Earth's surface has increased.

Most scientists think that this has been caused by an increase in the concentration of greenhouse gases in the atmosphere.

Give **one** reason why some scientists do **not** accept this theory.

(1)

(Total 8 marks)

Q2.

This question is about the fractions obtained from crude oil.

- (a) Crude oil is separated into fractions by fractional distillation.

The fractions obtained from crude oil include:

- lubricating oil
- naphtha
- petroleum gases.

Table 1 shows the boiling point range of these fractions.

Table 1

Fraction	Boiling point range in °C
Lubricating oil	300–350
Naphtha	90–200
Petroleum gases	< 25

Explain how these fractions are obtained from crude oil by fractional distillation.

(4)

- (b) Fractions from crude oil can be processed to produce feedstock for the petrochemical industry.

Which **two** are useful materials produced from this feedstock?

Tick (✓) **two** boxes.

Alloys

Ceramics

Detergents

Fertilisers

Solvents

(2)

Another fraction obtained from crude oil is petrol.

- (c) Petrol contains a hydrocarbon with the formula C_9H_{20}

Complete the equation for the complete combustion of C_9H_{20}

You should balance the equation.



(2)

- (d) Petrol obtained from crude oil contains sulfur impurities.

Explain why sulfur impurities are removed before petrol is burned in car engines.

(2)

- (e) **Table 2** shows information about two more fractions obtained from crude oil.

Table 2

Fraction	Range of number of carbon atoms in each molecule
Kerosene	11–15
Heavy fuel oil	20–40

A student predicted that heavy fuel oil is more viscous than kerosene.

The student's prediction was correct.

Justify the student's prediction.

(2)

The heavy fuel oil fraction can be processed to produce smaller hydrocarbon molecules.

- (f) Name the process which produces smaller hydrocarbon molecules from heavy fuel oil.

Give the conditions used in this process.

Name of process _____

Conditions _____

(3)

- (g) Hydrocarbon molecules containing seven and eight carbon atoms can be produced when heavy fuel oil is processed.

Which pair of hydrocarbon molecules would **both** turn bromine water colourless?

Tick (✓) **one** box.

C_7H_{14} and C_8H_{16}

C_7H_{14} and C_8H_{18}

C_7H_{16} and C_8H_{16}

C_7H_{16} and C_8H_{18}

(1)
(Total 16 marks)