

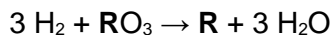
All questions are for separate science students only

**Q1.**

This question is about the extraction of metals.

Element **R** is extracted from its oxide by reduction with hydrogen.

The equation for the reaction is:



- (a) The sum of the relative formula masses ( $M_r$ ) of the reactants ( $3 \text{H}_2 + \text{RO}_3$ ) is 150

Calculate the relative atomic mass ( $A_r$ ) of **R**.

Relative atomic masses ( $A_r$ ): H = 1 O = 16

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Relative atomic mass ( $A_r$ ) of **R** = \_\_\_\_\_

(2)

- (b) Identify element **R**.

You should use:

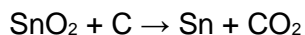
- your answer to part (a)
- the periodic table.

Identity of **R** = \_\_\_\_\_

(1)

- (c) Carbon is used to extract tin (Sn) from tin oxide ( $\text{SnO}_2$ ).

The equation for the reaction is:



Calculate the percentage atom economy for extracting tin in this reaction.

Relative atomic masses ( $A_r$ ): C = 12 O = 16 Sn = 119

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(4)  
(Total 10 marks)

**Q2.**

This question is about organic compounds.

- (a) Butane is an alkane with small molecules.

Complete the sentence.

Choose the answer from the box.

fertiliser	formulation	fuel
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Butane can be used as a \_\_\_\_\_.

(1)

- (b) Poly(propene) is a polymer.

What is the name of the monomer used to produce poly(propene)?

Tick (✓) **one** box.

Propane

Propanoic acid

Propanol

Propene

(1)

Ethene and steam react to produce ethanol.

The equation for the reversible reaction is:



- (c) The reaction produces a maximum theoretical mass of 400 kg of ethanol from 243 kg of ethene and 157 kg of steam.

A company produces 380 kg of ethanol from 243 kg of ethene and 157 kg of steam.

The percentage yield of ethanol is less than 100%

Calculate the percentage yield of ethanol.

Use the equation:

$$\text{percentage yield of ethanol} = \frac{\text{mass of ethanol actually made}}{\text{maximum theoretical mass of ethanol}} \times 100$$

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Percentage yield = \_\_\_\_\_ %

(2)

- (d) What are **two** possible reasons why the percentage yield of ethanol is less than 100%?

Tick (✓) **two** boxes.

Ethanol is the only product of the reaction.

Ethanol is very unreactive.

Some ethanol changes back into ethene and steam.

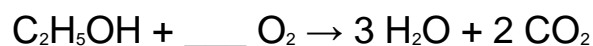
Some ethanol escapes from the apparatus.

Some ethanol reacts with steam.

(2)

- (e) Ethanol burns in oxygen.

Balance the equation for the reaction.



(1)

- (f) Two processes for producing ethanol are:

- fermentation
- hydration (reacting ethene with steam).

The table below shows information about the processes.

Feature	Process	
	Fermentation	Hydration

Raw material	sugar	crude oil
Energy usage	low	high
Rate of reaction	slow	fast
Purity of ethanol	15%	98%

Give **two** advantages and **two** disadvantages of using fermentation to produce ethanol.

Advantage of fermentation 1 \_\_\_\_\_

\_\_\_\_\_

Advantage of fermentation 2 \_\_\_\_\_

\_\_\_\_\_

Disadvantage of fermentation 1 \_\_\_\_\_

\_\_\_\_\_

Disadvantage of fermentation 2 \_\_\_\_\_

\_\_\_\_\_

(4)

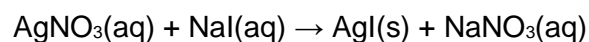
(Total 11 marks)

### Q3.

This question is about silver iodide.

Silver iodide is produced in the reaction between silver nitrate solution and sodium iodide solution.

The equation for the reaction is:



(a) A student investigated the law of conservation of mass.

This is the method used.

1. Pour silver nitrate solution into a beaker labelled **A**.
2. Pour sodium iodide solution into a beaker labelled **B**.
3. Measure the masses of both beakers and their contents.
4. Pour the solution from beaker **B** into beaker **A**.
5. Measure the masses of both beakers and their contents again.

The table below shows the student's results.

	Mass before mixing in g	Mass after mixing in g
Beaker <b>A</b> and contents	78.26	108.22
Beaker <b>B</b> and contents	78.50	48.54

Explain how the results demonstrate the law of conservation of mass.

You should use data from table above in your answer.

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(2)

- (b) Suggest how the student could separate the insoluble silver iodide from the mixture at the end of the reaction.

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(1)

The student purified the separated silver iodide.

This is the method used.

1. Rinse the silver iodide with distilled water.
2. Warm the silver iodide.

- (c) Suggest **one** impurity that was removed by rinsing with water.

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(1)

- (d) Suggest why the student warmed the silver iodide.

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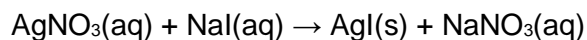


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(1)

- (e) Calculate the percentage atom economy for the production of silver iodide in this reaction.

The equation for the reaction is:



Give your answer to 3 significant figures.

Relative formula masses:

( $M_r$ ):  $\text{AgNO}_3 = 170$      $\text{NaI} = 150$      $\text{AgI} = 235$      $\text{NaNO}_3 = 85$

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Percentage atom economy (3 significant figures) = \_\_\_\_\_ %

(4)

- (f) Give **one** reason why reactions with a high atom economy are used in industry.

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(1)

(Total 10 marks)

#### Q4.

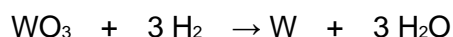
This question is about the extraction of metals.

- (a) Tungsten is a metal.

The symbol of tungsten is W

Tungsten is produced from tungsten oxide by reaction with hydrogen.

The equation for the reaction is:



Calculate the percentage atom economy when tungsten is produced in this reaction.

Use the equation:

$$\text{percentage atom economy} = \frac{184}{(M_r \text{ WO}_3) + (3 \times M_r \text{ H}_2)} \times 100$$

Relative formula masses ( $M_r$ ):  $\text{WO}_3 = 232$      $\text{H}_2 = 2$

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Percentage atom economy = \_\_\_\_\_%

(2)

Aluminium is extracted from aluminium oxide.

(b) 38% of a rock sample is aluminium oxide.

Calculate the mass of aluminium oxide in 40 kg of the rock sample.

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Mass of aluminium oxide = \_\_\_\_\_ kg

(2)

(c) The formula of aluminium oxide is  $\text{Al}_2\text{O}_3$

Calculate the relative formula mass ( $M_r$ ) of aluminium oxide.

Relative atomic masses ( $A_r$ ):       $\text{O} = 16$        $\text{Al} = 27$

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Relative formula mass ( $M_r$ ) = \_\_\_\_\_

(2)

(d) 60.0 kg of aluminium oxide produces a maximum of 31.8 kg of aluminium.

In an extraction process only 28.4 kg of aluminium is produced from 60.0 kg of aluminium oxide.

Calculate the percentage yield.

Give your answer to 3 significant figures.

Use the equation:

$$\text{percentage yield} = \frac{\text{mass of product actually made}}{\text{maximum theoretical mass of product}} \times 100$$

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Percentage yield = \_\_\_\_\_%

(3)

(e) Extracting metals by electrolysis is a very expensive process.

Explain why aluminium is extracted using electrolysis and not by reduction with carbon.

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(2)

(Total 11 marks)

**Q5.**

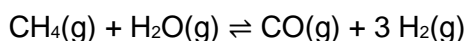
This question is about reversible reactions and equilibrium.

Hydrogen is used to produce ammonia in the Haber process.

The hydrogen is made in two stages.

**Stage 1** is the reaction of methane and steam to produce carbon monoxide and hydrogen.

The equation for the reaction is:



- (a) Calculate the atom economy for the formation of hydrogen in **stage 1**.

Relative atomic masses ( $A_r$ ): H = 1 C = 12 O = 16

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Atom economy = \_\_\_\_\_%

(2)

- (b) Explain why a low pressure is used in **stage 1**.

Give your answer in terms of equilibrium.

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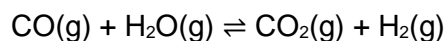
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(2)

- (c) **Stage 2** uses the carbon monoxide produced in **stage 1**.

The carbon monoxide is reacted with more steam to produce carbon dioxide and more hydrogen.

The equation for the reaction in **stage 2** is:



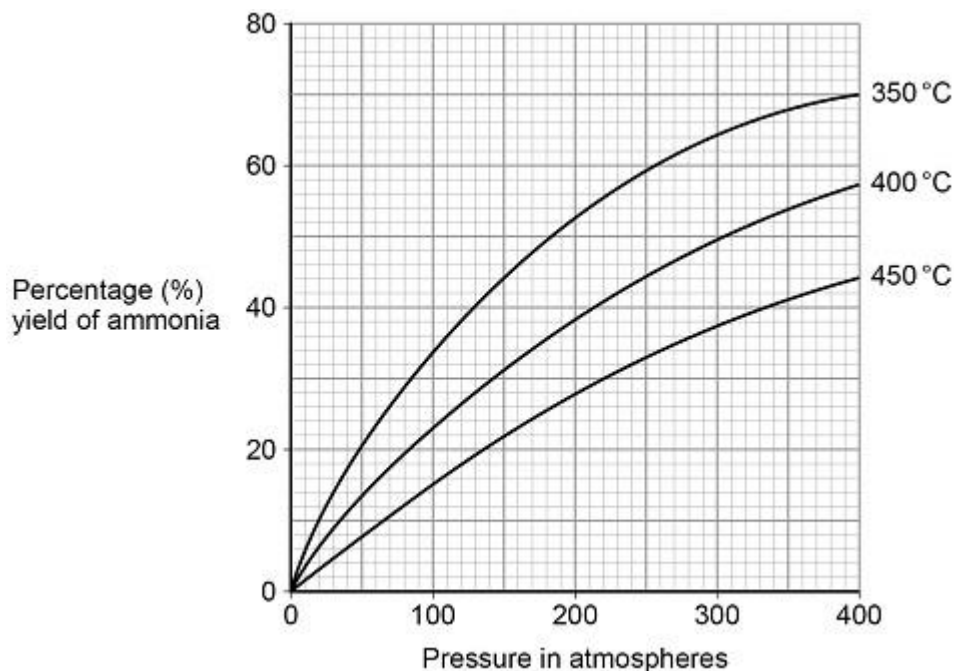
What is the effect of increasing the pressure on the equilibrium yield of hydrogen in **stage 2**?

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(1)

The graph below shows the percentage yield of ammonia produced at different temperatures and pressures in the Haber process.



A temperature of 450 °C and a pressure of 200 atmospheres are used in the Haber process.

- (d) A student suggested that a temperature of 350 °C and a pressure of 285 atmospheres could be used instead of those used in the Haber process.

Determine how many times greater the percentage yield of ammonia obtained would be.

Use the graph.

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Percentage yield = \_\_\_\_\_ times greater

(3)

- (e) A pressure of 285 atmospheres is **not** used in the Haber process instead of 200 atmospheres.

Give **one** reason why.

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(1)

- (f) How does the graph above show that the forward reaction in the Haber process is exothermic?

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(1)

- (g) World production of ammonia is now about 30 times greater than it was in 1950.

Suggest why the demand for ammonia has increased.

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(2)

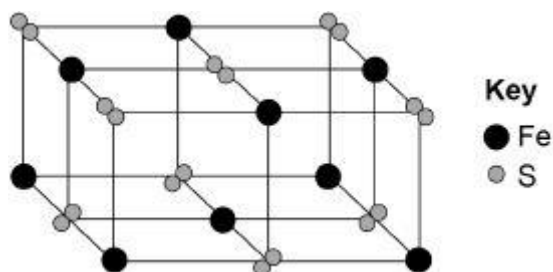
(Total 12 marks)

**Q6.**

This question is about metals and metal compounds.

- (a) Iron pyrites is an ionic compound.

The diagram below shows a structure for iron pyrites.



Determine the formula of iron pyrites.

Use the diagram above.

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(1)

- (b) An atom of iron is represented as  ${}_{26}^{56}\text{Fe}$

Give the number of protons, neutrons and electrons in this atom of iron.

Number of protons \_\_\_\_\_

Number of neutrons \_\_\_\_\_

Number of electrons \_\_\_\_\_

(3)

(c) Iron is a transition metal.

Sodium is a Group 1 metal.

Give **two** differences between the properties of iron and sodium.

1. \_\_\_\_\_

\_\_\_\_\_

2. \_\_\_\_\_

\_\_\_\_\_

(2)

Nickel is extracted from nickel oxide by reduction with carbon.

(d) Explain why carbon can be used to extract nickel from nickel oxide.

\_\_\_\_\_

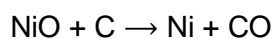
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\_\_\_\_\_

(2)

(e) An equation for the reaction is:



Calculate the percentage atom economy for the reaction to produce nickel.

Relative atomic masses ( $A_r$ ): C = 12 Ni = 59

Relative formula mass ( $M_r$ ): NiO = 75

Give your answer to 3 significant figures.

\_\_\_\_\_

\_\_\_\_\_

\_\_\_\_\_

\_\_\_\_\_

\_\_\_\_\_

Percentage atom economy = \_\_\_\_\_ %

(3)

(Total 11 marks)

**Q7.**

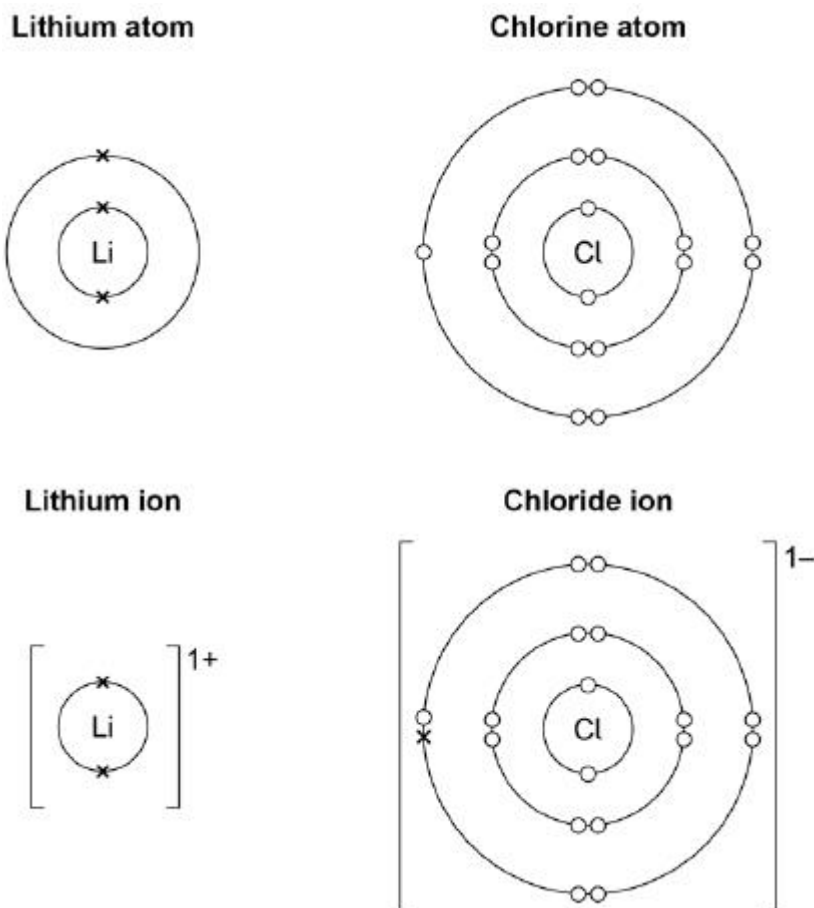
This question is about metal compounds.

- (a) Lithium reacts with chlorine to produce lithium chloride.

When lithium atoms and chlorine atoms react to produce lithium chloride, lithium ions and chloride ions are formed.

The diagram shows the electronic structures of the atoms and ions.

The symbols **o** and **x** are used to represent electrons.



Describe what happens when a lithium atom reacts with a chlorine atom.

Answer in terms of electrons.

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(4)

Zinc sulfate can be made by two methods.

The equations for the two methods are:



- (b) Calculate the percentage atom economy for making zinc sulfate in **Method 1**.

Use the equation:

percentage atom economy =

$$\frac{\text{relative formula mass of ZnSO}_4}{\text{relative formula mass of ZnO} + \text{relative formula mass of H}_2\text{SO}_4} \times 100$$

Give your answer to 3 significant figures.

Relative formula masses ( $M_r$ ): ZnO = 81 H<sub>2</sub>SO<sub>4</sub> = 98 ZnSO<sub>4</sub> = 161

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Percentage atom economy = \_\_\_\_\_ %

(3)

- (c) **Method 1** gives a higher percentage atom economy for making zinc sulfate than **Method 2**.

Give a reason why it is important to use a reaction with a high atom economy.

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(1)

- (d) A student uses 50 cm<sup>3</sup> of a zinc sulfate solution of 80 g/dm<sup>3</sup>

What mass of zinc sulfate is dissolved in 50 cm<sup>3</sup> of this zinc sulfate solution?

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Mass = \_\_\_\_\_ g

(2)

(Total 10 marks)

**Q8.**

A scientist produces zinc iodide ( $\text{ZnI}_2$ ).

This is the method used.

1. Weigh 0.500 g of iodine.
2. Dissolve the iodine in ethanol.
3. Add an excess of zinc.
4. Stir the mixture until there is no further change.
5. Filter off the excess zinc.
6. Evaporate off the ethanol.

(a) Ethanol is flammable.

Suggest how the scientist could carry out **Step 6** safely.

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(1)

(b) Explain why the scientist adds excess zinc rather than excess iodine.

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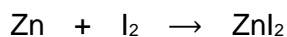
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(3)

(c) Calculate the minimum mass of zinc that needs to be added to 0.500 g of iodine so that the iodine fully reacts.

The equation for the reaction is:



Relative atomic masses ( $M_r$ ): Zn = 65    I = 127

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Minimum mass of zinc = \_\_\_\_\_ g

(3)

A different scientist makes zinc iodide by the same method.

The scientist obtains 12.5 g of zinc iodide.

The percentage yield in this reaction is 92.0%.

- (d) What is the maximum theoretical mass of zinc iodide produced in this reaction?

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Maximum theoretical mass = \_\_\_\_\_ g

(3)

- (e) Suggest **one** reason why the percentage yield in this reaction is **not** 100%.

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(1)

- (f) The scientist makes a solution of zinc iodide with a concentration of  $0.100 \text{ mol / dm}^3$

Calculate the mass of zinc iodide ( $\text{ZnI}_2$ ) required to make  $250 \text{ cm}^3$  of this solution.

Relative atomic masses ( $A_r$ ): Zn = 65 I = 127

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Mass = \_\_\_\_\_ g

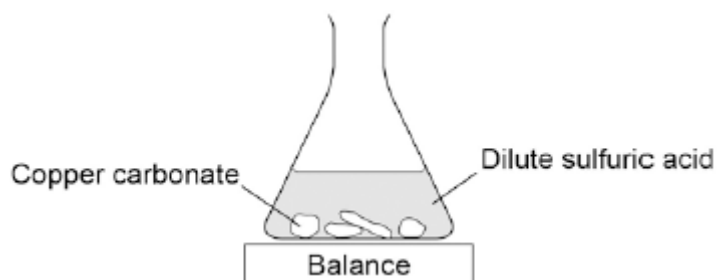
(3)

(Total 14 marks)

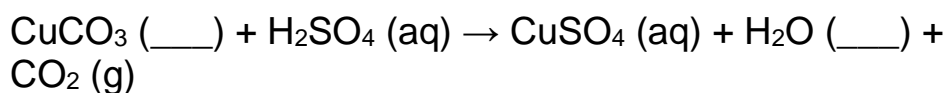
**Q9.**

A student investigated the reaction of copper carbonate with dilute sulfuric acid.

The student used the apparatus shown in the figure below.



- (a) Complete the state symbols in the equation.



(2)

- (b) Why did the balance reading decrease during the reaction?

Tick **one** box.

The copper carbonate broke down.

A salt was produced in the reaction.

A gas was lost from the flask.

Water was produced in the reaction.



- (e) Give **one** reason why is it important for the percentage atom economy of a reaction to be as high as possible.

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(1)

(Total 13 marks)

**Q10.**

A student investigated the reactions of copper carbonate and copper oxide with dilute hydrochloric acid.

In both reactions one of the products is copper chloride.

- (a) Describe how a sample of copper chloride crystals could be made from copper carbonate and dilute hydrochloric acid.

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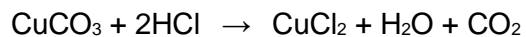
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(4)

- (b) A student wanted to make 11.0 g of copper chloride.

The equation for the reaction is:



Relative atomic masses,  $A_r$ : H = 1; C = 12; O = 16; Cl = 35.5; Cu = 63.5

Calculate the mass of copper carbonate the student should react with dilute hydrochloric acid to make 11.0 g of copper chloride.

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Mass of copper carbonate = \_\_\_\_\_ g (4)

- (c) The percentage yield of copper chloride was 79.1 %.

Calculate the mass of copper chloride the student actually produced.

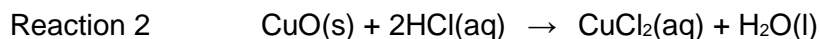
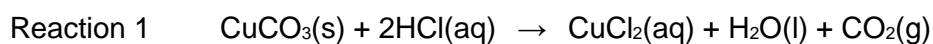
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Actual mass of copper chloride produced = \_\_\_\_\_ g (2)

- (d) Look at the equations for the two reactions:



Reactive formula masses: CuO = 79.5; HCl = 36.5; CuCl<sub>2</sub> = 134.5; H<sub>2</sub>O = 18

The percentage atom economy for a reaction is calculated using:

$$\frac{\text{Relative formula mass of desired product from equation}}{\text{Sum of relative formula masses of all reactants from equation}} \times 100$$

Calculate the percentage atom economy for Reaction 2.

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Percentage atom economy = \_\_\_\_\_ % (3)

- (e) The atom economy for Reaction 1 is 68.45 %.  
Compare the atom economies of the two reactions for making copper chloride.

Give a reason for the difference.

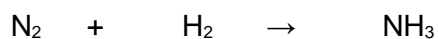
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(1)  
(Total 14 marks)

**Q11.**

- (a) Nitrogen and hydrogen are passed over iron to produce ammonia in the Haber Process.

Balance the equation for the reaction.



(1)

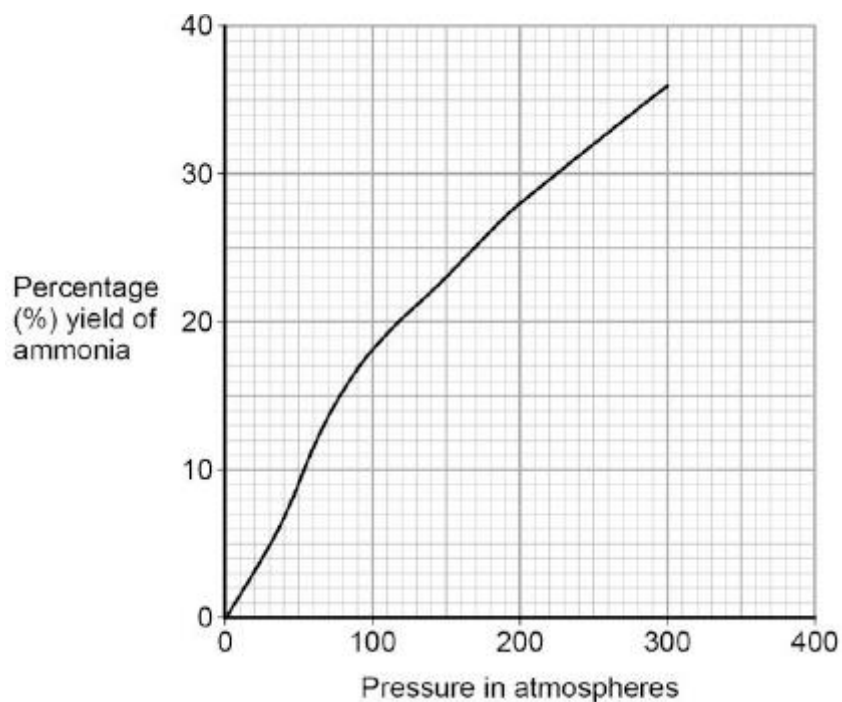
- (b) What is iron used for in the Haber process?

Tick **one** box.

- |          |                          |
|----------|--------------------------|
| catalyst | <input type="checkbox"/> |
| fuel     | <input type="checkbox"/> |
| monomer  | <input type="checkbox"/> |
| reactant | <input type="checkbox"/> |

(1)

- (c) The figure below shows how the percentage yield of ammonia changes with pressure.



Describe the trend shown in the figure above.

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**(1)**

- (d) Use the figure above to determine the difference in percentage yield of ammonia at 150 atmospheres pressure and 250 atmospheres pressure.

Difference in percentage yield of ammonia = \_\_\_\_\_ %

**(2)**

**(Total 5 marks)**